

Chemistry Terms Study Sheet

1. acid A substance that donates protons in a solution.
2. acid rain Precipitation that is more acidic than normal due to air pollution.
3. alloy A mixture of metals.
4. alpha particle A type of radiation consisting of two protons and two neutrons.
5. anode The electrode where oxidation occurs.
6. aqueous A solution where the solvent is water.
7. atom The smallest unit of an element.
8. base A substance that accepts protons in a solution.
9. battery A device that stores and provides electrical energy.
10. beta particle A type of radiation consisting of high-energy electrons.
11. biofuel Fuel derived from biological sources like plants.
12. biodegradable Capable of being broken down by natural processes.
13. boiling The process where a liquid changes into a gas.
14. buffer A solution that resists changes in pH.
15. catalyst Substance that speeds up a chemical reaction without being consumed.
16. cathode The electrode where reduction occurs.
17. chemical bond The force that holds atoms together in a molecule.
18. chemical reaction Process where substances are transformed into different substances.
19. chromatography A technique for separating mixtures based on differences in their movement through a medium.
20. colloid A mixture where very small particles of one substance are evenly distributed throughout another.
21. combustion A chemical reaction that involves the burning of a substance in oxygen.
22. compound A substance formed from two or more elements chemically united.
23. concentration The amount of solute in a given amount of solvent.
24. condensation The process where a gas changes into a liquid.
25. corrosion The gradual destruction of metals by chemical reactions with the environment.
26. covalent Type of bond where atoms share electrons.
27. decomposition A chemical reaction where a single compound breaks down into simpler substances.

28. deposition	The process where a gas changes directly into a solid.
29. dilution	Adding more solvent to decrease the concentration of a solution.
30. diffusion	The spreading of particles from an area of higher concentration to an area of lower concentration.
31. dipole	A molecule with two poles of opposite charge.
32. distillation	A process used to separate mixtures based on differences in boiling points.
33. electron	A negatively charged particle orbiting the nucleus.
34. electrolysis	A process that uses electricity to drive a chemical reaction.
35. electrolyte	A substance that conducts electricity when dissolved in water.
36. electroplating	The process of coating a metal object with a thin layer of another metal using electrolysis.
37. element	A substance that cannot be broken down into simpler substances.
38. emulsification	The process of mixing two immiscible liquids.
39. endothermic	A reaction that absorbs energy.
40. equation	A statement that shows the equality of two expressions.
41. equilibrium	The state where the rates of the forward and reverse reactions are equal.
42. exothermic	A reaction that releases energy.
43. fermentation	The chemical breakdown of a substance by bacteria, yeasts, or other microorganisms.
44. filtration	A process used to separate solids from liquids using a filter.
45. fission	The splitting of a nucleus into smaller parts.
46. formula	A representation of the elements in a compound.
47. fossil fuel	Nonrenewable energy sources like coal, oil, and natural gas.
48. freezing	The process where a liquid changes into a solid.
49. fusion	The combining of nuclei to form a heavier nucleus.
50. galvanic	A type of electrochemical cell that generates electrical energy.
51. gas	A state of matter with no fixed shape or volume.
52. geothermal	Energy derived from the heat within the earth.
53. greenhouse gas	Gases that trap heat in the atmosphere.
54. half-life	The time it takes for half of a radioactive substance to decay.
55. hydrogen bond	A strong type of intermolecular force involving hydrogen.
56. hydrolysis	A chemical reaction involving water.

57. indicator	A substance that changes color to indicate the presence of an acid or base.
58. intermolecular	Forces between molecules.
59. intramolecular	Forces within a molecule.
60. ion	An atom or molecule with a net electric charge.
61. ionic	Type of bond where atoms transfer electrons.
62. isotope	Atoms of the same element with different numbers of neutrons.
63. kinetics	The study of the rate of chemical reactions.
64. lattice	A regular arrangement of particles in a crystalline solid.
65. liquid	A state of matter with a definite volume but no fixed shape.
66. melting	The process where a solid changes into a liquid.
67. metallic	Type of bond found in metals, involving a sea of electrons.
68. methane	A potent greenhouse gas produced by livestock and decay of organic matter.
69. mixture	A combination of two or more substances that are not chemically combined.
70. molarity	The number of moles of solute per liter of solution.
71. mole	The unit for amount of substance, containing Avogadro's number of particles.
72. molecule	Two or more atoms bonded together.
73. monomer	The small building block that makes up polymers.
74. neutron	A neutral particle in an atom's nucleus.
75. neutron star	A dense celestial object composed mainly of neutrons.
76. nonrenewable	Resources that cannot be replenished in a short period of time.
77. nuclear energy	Energy released during nuclear reactions.
78. oxidation	The loss of electrons in a reaction.
79. ozone	A molecule composed of three oxygen atoms, important in the atmosphere.
80. periodic table	A chart of the elements arranged by atomic number.
81. pH	A measure of how acidic or basic a solution is.
82. pH scale	A measure of acidity or alkalinity of a solution.
83. photosynthesis	The process by which plants convert sunlight into chemical energy.
84. plasma	A state of matter similar to gas but consisting of charged particles.
85. plastic	A type of synthetic polymer.

86. polymer	A large molecule made up of repeating units.
87. precipitate	A solid that forms and settles out of a liquid mixture.
88. products	The substances formed in a chemical reaction.
89. proton	A positively charged particle in an atom's nucleus.
90. radioactivity	The emission of particles or energy from unstable nuclei.
91. reactants	The starting substances in a chemical reaction.
92. reduction	The gain of electrons in a reaction.
93. renewable	Resources that can be replenished naturally.
94. respiration	The process by which cells convert glucose and oxygen into energy.
95. redox	A type of reaction involving the transfer of electrons.
96. solid	A state of matter with a definite shape and volume.
97. solubility	The ability of a substance to dissolve in a solvent.
98. solute	The substance that is dissolved in a solution.
99. solution	A homogeneous mixture of two or more substances.
100. solvent	The substance that dissolves the solute in a solution.
101. stoichiometry	The calculation of reactants and products in chemical reactions.
102. sublimation	The process where a solid changes directly into a gas.
103. suspension	A mixture in which particles are dispersed throughout the bulk of a fluid.
104. synthesis	A chemical reaction in which two or more substances combine to form a new compound.
105. thermodynamics	The study of energy changes in chemical reactions.
106. titration	A technique to determine the concentration of a substance in a solution.
107. valence	The outermost electrons of an atom.
108. van der Waals	Weak intermolecular forces between molecules.
109. voltaic	Another name for a galvanic cell.
110. wind energy	Energy harnessed from the wind.